Studying the Photodegradation of Congo Red Dye from Aqueous Solutions Using Bimetallic Au–Pd/TiO$_2$ Photocatalyst

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Abstract: In this study, the photodegradation of Congo red dye (CR) in aqueous solution was investigated using Au-Pd/TiO$_2$ as photocatalyst. The concentration of dye, dosage of photocatalyst, amount of H$_2$O$_2$, pH of the medium and temperature were examined to find the optimum values of these parameters. It has been found that 28 ppm was the best dye concentration. The optimum amount of photocatalyst was 0.09 g/75 mL of dye solution when the degradation percent was ~ 96 % after irradiation time of 12 hours, while the best amount of hydrogen peroxide was 7μl/75 mL of dye solution at degradation percent ~97 % after irradiation time of 10 hours, whereas pH 5 was the best value to carry out the reaction at the highest degradation percent. In addition, temperature tested at range of (25-55) °C, and it has been figured out which photodegradation percent of dye increase with raising temperature (degradation percent was ~ 98% after irradiation time of 4 hours at 55 °C), and the activation energy of the reaction was calculated (34.8016 kJ/mole) from Arrhenius law. The thermodynamic functions ΔH°, ΔG°, and ΔS° were obtained, where ΔH° and ΔG° are positive value which means that the reaction is endothermic and non-spontaneous respectively, while ΔS° has a negative value, thus indicates that the reactants are more disordered than the excited intermediate formed. The kinetic of the reaction was studied, and it has been found that the photocatalytic reaction follows pseudo first order reaction.

Keywords: Congo red, Kinetic, Photodegradation, Thermodynamic, Titanium dioxide

Introduction: Many studies have been extensively investigated the heterogeneous photocatalysis oxidations, which performed with light irradiated semiconductors dispersions due to their highly efficiency as an efficient rout for degradation of toxic organic and inorganic compounds to carbon dioxide and water (1). Among the semiconductors, titanium dioxide TiO$_2$ is considered as an important photocatalyst, because of intense catalytic activity, when is illuminated with photons, whose energy equal to band-gap energy thus creation of electron-hole pairs on the surface of TiO$_2$ (2). Excited-state electrons and holes can recombine and dissipate the input energy as heat, or react with adsorbed compounds on the surface of the TiO$_2$. After the reaction of the electron-hole pairs with mentioned compounds, OH radicals, O$_2$, H$_2$O$_2$ can be produced, which play important roles in the photocatalytic reaction (3). TiO$_2$ exhibits wide band gap of around 3.2 eV, which limits its absorption in the visible area, therefore, many researchers have used doping method to enhance the photocatalytic performance. Thus, the optical absorption range are extended from ultraviolet to visible light, that attributed to generate the new energy level in the band gap of TiO$_2$ due to the distribution of nanoparticles of metal in the TiO$_2$ (4-8). It has been found that the treatment of dye effluents is not easy because of their various synthetic origins and aromatic structure. Congo red Fig. 1 is one of the most common synthetic dyes used in the textile manufacturing, it belongs to azo dyes, which are known to be non-biodegradable, because the existence of one or more azo groups (-N=N-) in its structure (9). The release of these azo compounds into water is undesirable, and it can be toxic to the living organisms. Degradation of naphthalene and benzene rings in Congo red cannot be applied with ordinary methods. Because of its aromatic structure, Congo red possesses high physiochemical, thermal
and optical stability (10, 11). The absorption spectrum of the Congo red Fig. 2 shows two peaks, (334 nm and 497 nm). The absorption band at 344 nm is associated with the π–π* transition of the aromatic ring and the band at 497 nm can be corresponding to the n–π* transition of the lone pair present in the N atom of the azo groups (–N=N–) (12). Different studies have investigated the mineralization of Congo red dye using different methods; photocatalytic degradation (13, 14), biodegradation (15), adsorption (16). In this study, Au-Pd/TiO₂ (17) was used as photocatalyst to degrade the Congo red dye with studying the effect of various parameters; concentration of dye, dosage of photocatalyst, amount of H₂O₂, pH of the medium and the effect of temperature.

![Figure 1. Molecular structure of Congo red dye (CR)](image)

Figure 1. Molecular structure of Congo red dye (CR)

![Figure 2. Absorption spectrum of Congo red dye](image)

Figure 2. Absorption spectrum of Congo red dye

**Materials and Methods**

**Materials**

Congo red (CR), (C₃₂H₂₂N₆Na₂O₆S₂) sodium salt of benzidine diazo-bis-1- naphthylamine-4-sulfonic acid was supplied from Reidel-De Haen (Germany), Au-Pd/TiO₂ from Dr. Sanaa Tareq (17). Hydrogen peroxide (H₂O₂) was from Carlo Erba company at purity 30%. Hydrochloric acid was provided from BDH and sodium hydroxide was supplied from Riedel-De Haen AG Seelze-Hannover.

**Irradiation System**

In this work, the irradiation system consists of Pyrex photoreaction cell 75 mL capacity with a quartz window fitted with a focusing lens to ensure parallel beam of light. The mercury lamp 125 watt is placed 10 cm apart from the photoreaction cell. The cell is supplied with two openings of 0.5 cm in diameter and one of these is used for sampling processes. The magnetic stirrer (wise) from Daihan scientific is used to keep the catalyst aqueous suspension in homogeneous form during the photolysis experiments, and the circulating thermostat (Recirculationg chiller) from Daihan scientific is used to control the reaction temperature. Absorption values were recorded with SmartSpec™ plus spectrophotometer (BIO-RAD). All of samples were centrifuged using (centrifuge model 800 electric centrifuge), and photocatalysts were weighed using analytic sensitive balance (Sartorius) from Germany.

![Figure 3. Irradiation System; 1- Circulator, 2-Magnetic stirrer, 3- Photoreaction cell, 4-Lens, and 5- UV- source](image)

**Standard Calibration Curve of Congo Red Dye**

Figure 4 shows the standard calibration curve of Congo red of 497 nm, which was used to convert absorbance values of dye (before and after irradiation) to concentration according to beer-lambert law.

![Figure 4. Calibration curve of Congo red of 497 nm](image)

Figure 4. Calibration curve of Congo red of 497 nm

\[ y = 0.0534x + 0.0623 \]

\[ R^2 = 0.9983 \]
Procedure of the Degradation of Congo Red Dye

All of the experiments have been carried out in the photoreaction cell using (irradiation system) Fig. 3. Firstly, several concentrations of photocatalyst [Au-Pd/TiO$_2$], (0.03-0.09g) were added. Circulating thermostat has been used to control the reaction temperature, while magnetic stirrer has been used to mix the mixture. The solutions were irradiated with medium mercury lamp (125 watt). A sample of 5mL of the mixture from each experiment is taken at various time intervals (1 hour), and then the photocatalyst is removed from the samples by using centrifuge. After that, observing the photolysis process has been controlled by following up the absorbance of the Congo red dye (CR) at $\lambda_{max}$ 497 nm using UV-visible spectrophotometer. A series of experiments is repeated using different pH values (5-11), and the pH of the solution has been adjusted using hydrochloric acid (0.1M) and sodium hydroxide (0.1M). In addition, some experiments are carried out at various temperatures (25, 35, 45, and 55) °C to determine the best media and temperature for degradation of dye at optimum values of catalysts. The photodegradation percentage of Congo red dye is calculated using the following equation (18):

$$%D = \frac{C_0 - C_t}{C_0} \times 100$$  (1)

where, %D: The percentage degradation, $C_0$: The initial concentration (ppm) at zero time. $C_t$: The final concentration (mg/L) at any time.

Results and Discussion

Effect of Dye Concentration

Various concentrations of Congo red dye were irradiated under light of mercury lamp (125 watt) at irradiation time of an hour to examine the best concentration of dye. From Fig. 5, it is evident that the best degradation percent was at 28 ppm, which means the penetration of light was the best at this concentration (12), for this reason, 28 ppm became the optimum value to use through this study.

![Figure 5. The relation between % D of Congo red dye solution and various concentrations.](image)

Effect of Catalyst Dosage

Photocatalytic process of Congo red dye (CR) with different dosage of Au-Pd/TiO$_2$ catalyst was investigated. A series of experiments were carried out to assess the optimum catalyst loading by varying the amount of catalyst (Au-Pd/TiO$_2$) from 0.03-0.1g in aqueous solution of CR (28 ppm) at 25 °C. The percentage of degradation of dye (%D) has shown appreciable results, where Fig. 6 shows the highest degradation percentage of dye (~96%), which was at 0.09 g after 12 hours of irradiation time. The photodegradation percentage of the CR was found to speed up with increase in the catalyst loading, where on the surface of the photocatalyst (Au-Pd/TiO$_2$), the formed pair of positive hole (h$^+$) at valance band (VB) and electron (e$^-$) at conductive band (CB) will react with adsorbed water molecules, dissolved oxygen, and hydroxyl groups of the surface, producing free radicals of hydroxyl and superoxide radicals. These photogenerated radicals will attack the dye molecules, which are adsorbed on the photocatalyst surface forming intermediate compounds, which break into CO$_2$, H$_2$O, and ions of NO$_3^-$ and NH$_4^+$ as expressed in the following equations

$$\text{Au} - \text{Pd/TiO}_2 + h\nu \rightarrow \text{Au} - \text{Pd/TiO}_2\left(h^+_{\text{VB}} + e^-_{\text{CB}}\right)$$  (2)

$$\text{Au} - \text{Pd/TiO}_2\left(h^+_{\text{VB}}\right) + \text{H}_2\text{O} \rightarrow \text{H}^+ + \text{OH}^-$$  (3)

$$\text{Au} - \text{Pd/TiO}_2\left(e^-_{\text{CB}}\right) + \text{O}_2 \rightarrow \text{Au} - \text{Pd/TiO}_2 + \text{O}_2^*$$  (4)

CR molecules + OH$^+$ + O$_2^*$ → intermediate compounds

$$\rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{NO}_3^- + \text{NH}_4^+$$  (5)

It has been observed that the increase of the photocatalyst dosage more than the optimum value causes a decrease in the photodegradation percentage of CR dye, due to the reduce of the light penetration through the solution (11), and to the decrease in the number of active site on the surface of the catalyst resulted from the aggregation of Au-Pd/TiO$_2$ at high concentration. Our results are in good agreement with the literatures (11, 19).
Effect of H$_2$O$_2$

The effect of adding H$_2$O$_2$ on the photodegradation of Congo red dye (CR) was studied in the range of 3-9 μL/75 mL of the aqueous solution of CR at concentration of 28 ppm and using optimum value of photocatalyst (Au-Pd/TiO$_2$, 0.09 g/75 mL of solution) at 25°C. The results revealed that the optimum addition was 7 μL/75 mL of dye solution; where the photodegradation percent was (~ 97%) after 10 hours of irradiation time. Figure 7 shows the experimental results of adding H$_2$O$_2$ to the photocatalytic reaction. It has been observed that degradation percentage of dye increased with increase the amount of H$_2$O$_2$, which can be explained that the amount of oxygen is an important factor in the photocatalytic reaction. The oxygen is required as an electron scavenger to keep the reaction, so the oxygen flow into the photocatalytic reaction should be well regulated, as poor flow of oxygen can cause in the decrease of the efficiency of the photocatalytic reaction. For these reasons, H$_2$O$_2$ add to the photocatalytic reaction to generate more hydroxyl radicals, the effect can be expressed as in the following equations (20, 21):

\[
\begin{align*}
  \text{H}_2\text{O}_2 + \text{O}_2^* & \rightarrow \text{OH}^+ + \text{OH}^- + \text{H}_2\text{O} \\
  \text{H}_2\text{O}_2 + \text{hv} & \rightarrow 2\text{OH}^* \\
  \text{H}_2\text{O}_2 + \epsilon_{\text{CB}} & \rightarrow \text{OH}^* + \text{OH}^-
\end{align*}
\]

The photodegradation of CR dye decreased at higher amount of H$_2$O$_2$ beyond the optimum value, which can be attributed to hydroxyl radicals OH$^-$ and valence band holes could be consumed by H$_2$O$_2$, and radical-radical recombination as a competitive reaction can be taken in account, as described in the following equations:

\[
\begin{align*}
  \text{OH}^+ + \text{H}_2\text{O}_2 & \rightarrow \text{H}_2\text{O}_3^- + \text{H}_2\text{O} \\
  \text{OH}^+ + \text{H}_2\text{O}_2 & \rightarrow \text{O}_2 + \text{H}_2\text{O} \\
  \text{Au-Pd/TiO}_2(\text{hv}) + \text{H}_2\text{O}_2 & \rightarrow \text{Au-Pd/TiO}_2 + \text{O}_2 + 2\text{H}^+ \\
  \text{OH}^+ + \text{OH}^- & \rightarrow \text{H}_2\text{O}_2
\end{align*}
\]
Figure 8. The relation between % D of Congo red dye solution and irradiation time at different pH

Effect of Temperature

The effect of temperature has been investigated in the range of (25-55°C) in the aqueous solution of CR dye at the concentration 28 ppm and using optimum values of catalyst Au-Pd/TiO2 (0.09 g/75 mL of dye solution) and (7μL/75mL of dye solution) of H2O2. From Fig. 9, it can be noticed that the photodegradation of the dye increased with increasing temperature, as the photodegradation percentage at 55°C was (~ 98%) after irradiation time of 4 hours, when the colour of the dye disappeared completely. It has been reported that when the temperature of the photoreaction raised above of 80°C, it will enhanced the recombination of (electron-hole) and desorption reaction of the adsorbed compounds on the surface of photocatalyst leading to decrease the photocatalytic activity (26).

Figure 9. The relation between % D of Congo red dye solution and temperature

The activation energy Ea was calculated from the Arrhenius equation (27):

\[ \ln k = -\frac{E_a}{R} \left( \frac{1}{T} \right) + \ln A \quad (13) \]

where k is the rate constant of the photoreaction, Ea is the activation energy of the reaction, R is the gas constant, T is the absolute temperature in kelvin, and A is the frequency factor. Figure 10 shows the plot of ln k against 1/T where the slope of linear relationship is equal to –Ea/R. The entropy of activation ΔS°, the enthalpy of activation ΔH° and the free energy of activation ΔG° were calculated (Table 1) from the following equations (28)

\[ \ln A = \ln (kT/h) + \frac{\Delta S^o}{R} \quad (14) \]

\[ \Delta H^o = E_a - RT \quad (15) \]

\[ \Delta G^o = \Delta H^o - T\Delta S^o \quad (16) \]

From Table 1, It can be observed that the resulted activation energy of the photoreaction is very low (Ea = 34.8016 kJ mol⁻¹), which can be called the apparent activation as the true activation energy is nil. Because of that the photocatalytic reaction do not require the heating but it is accomplished by light leading to production of the excited (electron–hole) pair which consider the essential species for complete the reaction (29). The positive values of ΔH° and ΔG° refer to endothermic reaction and the reaction is non-spontaneous, respectively. Negative entropy was obtained which indicates that the reactants are more disordered than the formed intermediate (excited complex) (30).

Figure 10. Plot ln k against 1/T using Au-Pd/TiO2 as catalyst with Congo red dye

Table 1. Values of kinetics and thermodynamic parameters for the photocatalytic degradation of Congo red dye by using Au-Pd/TiO2 catalyst

<table>
<thead>
<tr>
<th>T (K)</th>
<th>k (h⁻¹)</th>
<th>Ea (kJ mol⁻¹)</th>
<th>ΔH° (kJ mol⁻¹)</th>
<th>ΔS° (kJ mol⁻¹)</th>
<th>ΔG° (kJ mol⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td>298</td>
<td>0.3097</td>
<td>32.3240</td>
<td>32.2408</td>
<td>-</td>
<td>74.0380</td>
</tr>
<tr>
<td>308</td>
<td>0.3705</td>
<td>34.8016</td>
<td>32.2408</td>
<td>0.1399</td>
<td>75.3547</td>
</tr>
<tr>
<td>318</td>
<td>0.7102</td>
<td>32.1577</td>
<td>32.2408</td>
<td>76.6714</td>
<td></td>
</tr>
<tr>
<td>328</td>
<td>1.0450</td>
<td>32.0746</td>
<td>77.9880</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

Table 1. Values of kinetics and thermodynamic parameters for the photocatalytic degradation of Congo red dye by using Au-Pd/TiO2 catalyst

\[ y = -185.9x + 12.775 \]

\[ R^2 = 0.9538 \]
Kinetics

A monomolecular Langmuir-Hinshelwood (L-H) modal is usually considered to quantify the heterogeneous photodegradation activity, where an adsorbed substrate is consumed at an initial rate given as shown in equation 17 (31):

\[-\frac{dC}{dt} = \frac{kKC}{1 + KC}\]  

where k is the specific rate constant, that change during photocatalytic reaction and K is the thermodynamic adsorption constant. In photocatalytic reaction, the photocatalyst concentration remains unchanged, so the rate of the reaction depends on the concentration of dye and the term pseudo can be used to prefix the rate of reaction. When the dye concentration is very high, so the surface of photocatalyst is saturated; a pseudo zero order reaction is assumed where, the rate constant of the reaction is independent on the concentration of the dye, as shown in the following equation (32):

\[-\frac{dC}{dt} = k_1 C\]  

While at low concentration of the dye; the photodegradation rate is dependent on the change in the dye concentration, which means the driving force of degradation is proportional to the dye concentration, in this case, equation 18 can be expressed in the following equation (33);

\[-\frac{dC}{dt} = k_2 C\]  

Table 2. Values of $R^2$ and rate constant for photodegradation of Congo red dye from plotting of relationships of first, second and zero orders at optimum values.

<table>
<thead>
<tr>
<th>Optimum values</th>
<th>First order reaction (plot ln C against t)</th>
<th>Second order reaction (plot 1/C against t)</th>
<th>Zero order reaction (plot (C₀-C) against t)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>$R^2$ k h⁻¹</td>
<td>$R^2$ k M h⁻¹</td>
<td>$R^2$ k M h⁻¹</td>
</tr>
<tr>
<td>0.09g of Au-Pd/TiO₂</td>
<td>0.9889 0.1384</td>
<td>0.9774 0.0127</td>
<td>0.5530 3.3655</td>
</tr>
<tr>
<td>7μL of H₂O₂</td>
<td>0.9169 0.2211</td>
<td>0.7565 0.0265</td>
<td>0.6733 3.3258</td>
</tr>
<tr>
<td>pH 5</td>
<td>0.8612 0.4380</td>
<td>0.4931 0.1866</td>
<td>0.2460 4.2100</td>
</tr>
<tr>
<td>55°C</td>
<td>0.9722 0.8694</td>
<td>0.7784 0.6494</td>
<td>-12.9700 9.1442</td>
</tr>
</tbody>
</table>

Conclusion:

To conclude, using of bimetallic Au-Pd/TiO₂ as a photocatalyst to break down the Congo red dye was efficient. The photodegradation was enhanced with testing different parameters such as; concentration of dye, dosage of the photocatalyst, amount of H₂O₂, pH of the solution and the temperature. It has been shown that the photodegradation percentages and rate of reaction increased respectively at the optimum values of these parameters. From calculated thermodynamic functions, it can be suggested that the reaction is endothermic and it is unspontaneous. By comparing the rate constant and regression coefficients, it can be indicated that the reaction follows pseudo first order.

Authors’ declaration:

- Conflicts of Interest: None.
- We hereby confirm that all the Figures and Tables in the manuscript are mine ours. Besides, the Figures and images, which are not mine ours, have been given the permission for republication attached with the manuscript.
- Ethical Clearance: The project was approved by the local ethical committee in University of Baghdad.

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دراسة التحلل الضوئي لصبغة الكونغو الحمراء في محاليلها المائية باستخدام العامل المساعد ثنائي الفلز ذهبي – بلاديوم / ثنائي أوكسيد التيتانيوم

سعداء عبد موسى
سناء طارق سناء
إيمان عدنان محمد

الخلاصة:

في هذه البحث، تم دراسة التحلل الضوئي لصبغة الكونغو الحمراء في محاليلها المائية باستخدام العامل المساعد ثنائي الفلز (Au-Pd/TiO$_2$). تم تحديد الكميات والظروف المثلى لأجراء التفاعل وتراكم الصبغة، كمية العامل المساعد، حامضية الوسط ودرجة الحرارة. ووجد أن أفضل تركيز لصبغة الكونغو الحمراء في التفاعل الضوئي كان عند 28 جزء بالمليون، في حين أن أفضل كمية للعامل المساعد كانت عند 0.09 غ/75 ملليتر من المحلول. كما أظهر الفحص أن التفاعل ضئوي الأكسدة سبب نسبة تحلل الصبغة تقريبا 96% بعد وقت تشعيع لمدة 12 ساعة.

أما بالنسبة للتفاعل التحلل الضوئي في محاليل الضوء، فقد أظهرت الفحص أن التفاعل ضئوي الأكسدة سبب نسبة تحلل الصبغة تقريبا 96% بعد وقت تشعيع لمدة 12 ساعة. وقد وجد أيضاً أن التفاعل الضوئي للصبغة يزداد مع زيادة درجة الحرارة عند اجراء التجارب في مدى حراري (25-55) درجة مئوية، حيث سجلت أعلى نسبة تفكك ضئوي للصبغة تقريبا 98% عند 55 درجة مئوية.

الكلمات المفتاحية: الدوال الثرموداينميكية، التفكك الضوئي، الحركة، صبغة الكونغو الحمراء، ثنائي أوكسيد التيتانيوم